

# COMPOSITIONAL, STRUCTURAL AND PHYSICAL STUDIES OF SOME GRAPHITE HEXAFLUOROARSENATES AND THEIR RELATIVES

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The F/As molar ratio for the vacuum-stable products obtained by treating graphite, at 20<sup>0</sup>, with (1) AsF<sub>5</sub>, (2) AsF<sub>5</sub> plus F<sub>2</sub>, and (3) O<sub>2</sub>AsF<sub>6</sub>, has been established to be 6. Route (1) gives mixtures of first and second-stage salts, whereas (2) and (3) readily yield first-stage material. This is attributed to the greater oxidizing potential of the reagents in (2) and (3). The volatiles from the route (1) synthesis are AsF<sub>5</sub> and AsF<sub>3</sub>, and from (2) AsF<sub>5</sub>. The interlayer spacing ( $I_c$ ) of the occupied graphite galleries in the vacuum-stable hexafluoroarsenates is ca. 7.6 Å and  $c \approx 7.6 + 3.35 (n-1)$  Å (where n is the stage). For the products of routes (1) and (2), prior to removal of volatiles,  $I_c \approx 8.0$  Å and  $c \approx 8.0 + 3.35(n-1)$  Å.

X-ray powder diffraction data for the first-stage salt C<sub>14</sub>AsF<sub>6</sub> have established that the AsF<sub>6</sub><sup>-</sup> species are nestled in contiguous three-fold sets of C-atom hexagons of the graphite. This requires a staggering of the enclosing C-atom layers, as in graphite. Each neighboring pair of C-atom sheets contains an ordered closest-packed assembly, appropriate for C<sub>14</sub>AsF<sub>6</sub>. Aside from the restriction imposed by AsF<sub>6</sub><sup>-</sup>-nestling, the layers of composition C<sub>14</sub>AsF<sub>6</sub> are otherwise randomly stacked. Materials prepared by direct interaction of graphite and AsF<sub>5</sub>, prior to removal of volatiles, have a larger interlayer spacing,  $I_c \approx 8.0$  Å than C<sub>14</sub>AsF<sub>6</sub> ( $I_c \approx 7.6$  Å). X-ray data for such materials are accounted for by a random distribution of un-nestled AsF<sub>x</sub> species, between eclipsed carbon-atom sheets.

The C<sub>x</sub>AsF<sub>6</sub> salts are all good conductors (approximately the conductivity of Al metal). Addition of fluorine to the salts decreases their conductivities substantially (approximately that of graphite).